

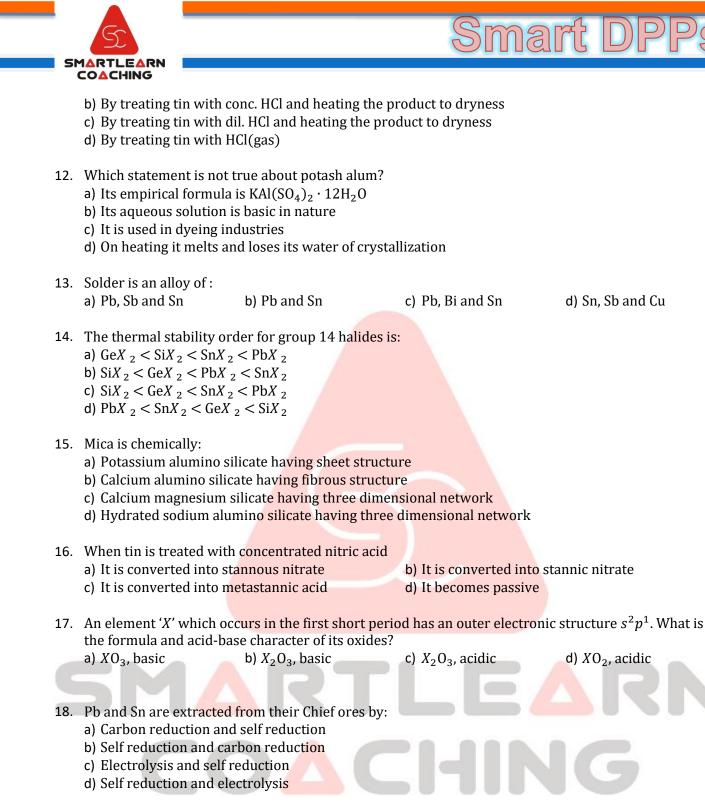


DPP				
DAILY PRACTICE PROBLEMS				

CLASS : XIth DATE :

SUBJECT : CHEMISTRY DPP No. : 2

Topic :- THE P-BLOCK ELEMENTS-1				
1.	Lead may be replaced from its salt solution by: a) Cu b) Au	c) Ag	d) Mg	
2.	Unstable lead compounds are a) $PbCl_4$, $PbBr_4$ and PbI_4 c) PbO , PbO_2 and Pb_3O_4	b) PbCl ₂ , PbBr ₂ and Pb d) PbCl ² ₄ ⁻ , PbCl ²⁻	b) PbCl ₂ , PbBr ₂ and PbI ₂ d) PbCl ₄ ²⁻ , PbCl ₆ ²⁻	
3.	Which acid is formed when SiF_4 reacts with waa) H_2SO_4 b) H_2SiF_4	ater? c) H ₂ SiF ₆	d) None of these	
4.	Which of the following reactions occurs at the cathode during the charging of lead accumulator? a) $Pb^{2+} + 2e \rightarrow Pb$ b) $Pb^{2+} + SO_4^{2-} \rightarrow PbSO_4$ c) $Pb \rightarrow Pb^{2+} + 2e$ d) $PbSO_4 + 2H_2O \rightarrow PbO_2 + 4H^+ + SO_4^{2-} + 2e$			
5.	The two type of bonds present in B ₂ H ₆ are cov a) Ionic b) Coordinate	alent and c) Hydrogen bridge	d) None of these	
6.	Which one shows most pronounced inert pair a) Si b) Sn	effect? c) Pb	d) C	
7.	Which of the following is an ore of lead? a) Galena b) Calamine	c) Malachite	d) Dolomite	
8. 9.	Soldiers of Napolean army while at Alps during freezing winter suffered a serious problem as regardsto the tin buttons of their uniforms. While metallic tin buttons got converted to grey powder. Thistransformation is related toa) An interaction with nitrogen of the air at very low to temperaturesb) A change in the partial pressure of oxygen in the airc) A change in the crystalline structure of tind) An interaction with water vapour contained in the humid airIn SiF_6^2 and SiCl_6^2 which one is known and why?a) SiF_6^2 because of small size of Fb) SiF_6^2 because of small size of Cl			
10.	Which of the following has structure similar to a) BN b) B	graphite? c) B ₄ C	d) B_2H_6	
11.	Tin(II) chloride (anhydrous) can be obtained : a) By melting tin in an atmosphere of Cl ₂			



19. Boron readily dissolves in:

- a) Conc. HCl
- b) Fused NaOH at 673 K
- c) Fused Na₂CO₃ at 1173K
- d) A mixture of conc. HNO_3 and conc. $H_2SO_4(1:2)$
- 20. The borax bead is chemically:
 - a) B_2O_3 b) $Na_2B_4O_7$ c) Na_3BO_3

d) $B_2O_3 + NaBO_2$