

DPP

DAILY PRACTICE PROBLEMS

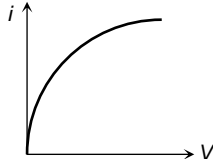
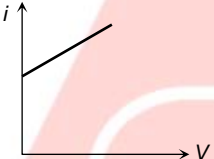
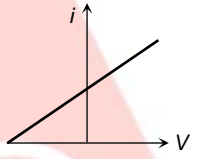
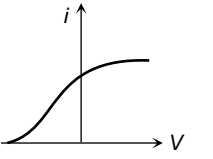
Class : XIIth

Date :

Subject : PHYSICS

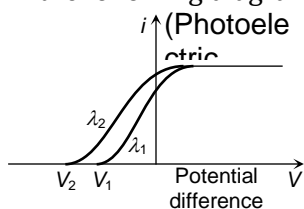
DPP No. : 2

Topic :- Dual nature of radiation and matter

- The uncertainty in the position of a particle is equal to the de-Broglie wavelength. The uncertainty in its momentum will be
 a) h/λ b) $2h/3\lambda$ c) λ/h d) $3\lambda/2h$
- The work functions for sodium and copper are $2eV$ and $4eV$. Which of them is suitable for a photocell with 4000 \AA light
 a) Copper b) Sodium c) Both d) Neither of them
- The curve between current (i) and potential difference (V) for a photo cell will be
 a)  b)  c)  d) 
- What will be the number of photons emitted per second by a 10 W sodium vapour lamp assuming that 90% of the consumed energy is converted into light? Wavelength of sodium light is 590 nm , $h = 6.63 \times 10^{-34} \text{ J}\cdot\text{s}$.
 a) 0.267×10^{18} b) 0.267×10^{19} c) 0.267×10^{20} d) 0.267×10^{17}
- For the Bohr's second orbit of circumference $2\pi r$, the de-Broglie wavelength of revolving electron will be
 a) $2\pi r$ b) πr c) $\frac{1}{2\pi r}$ d) $\frac{1}{4\pi r}$
- The work function of a metal is
 a) The energy for the electron to enter into the metal
 b) The energy for producing X-ray
 c) The energy is required for an electron to come out from metal surface
 d) None of these
- If the uncertainty in the position of proton is $6 \times 10^8 \text{ m}$, then the minimum uncertainty in its speed will be
 a) 1 cms^{-1} b) 1 ms^{-1} c) 1 mms^{-1} d) 100 ms^{-1}
- The work function for metals A, B and C are respectively 1.92 eV , 2.0 eV and 5 eV . According to Einstein's equation, the metals which will emit photo electrons for a radiation of wavelength 4100 \AA is/are
 a) None of these b) A only c) A and B only d) All the three metals
- Among the following four spectral regions, the photons has the highest energy in
 a) Infrared b) Violet c) Red d) Blue
- Kinetic energy of emitted cathode rays is dependent on
 a) Only voltage b) Only work function
 c) Both (a) and (b) d) It does not depend upon any physical quantity
- An electron is accelerated under a potential difference of 182 V . The maximum velocity of electron will be
 (Charge of an electron is $1.6 \times 10^{-19} \text{ C}$ and its mass is $9.1 \times 10^{-31} \text{ kg}$)
 a) $5.65 \times 10^6 \text{ m/s}$ b) $4 \times 10^6 \text{ m/s}$ c) $8 \times 10^6 \text{ m/s}$ d) $16 \times 10^6 \text{ m/s}$



12. If the voltage of X-rays tube is doubled, the intensity of X-rays will become
 - a) Half
 - b) Unchanged
 - c) Double
 - d) Four times
13. Bragg's law for X-rays is
 - a) $d \sin \theta = 2n\lambda$
 - b) $2d \sin \theta = n\lambda$
 - c) $n \sin \theta = 2\lambda d$
 - d) None of these
14. An electron of charge ' e ' coulomb passes through a potential difference of V volts. Its energy in 'joules' will be
 - a) V/e
 - b) eV
 - c) e/V
 - d) V
15. When cathode-rays strike a metal target of high melting point with a very high velocity, then which of the following are produced
 - a) α -rays
 - b) X-rays
 - c) Ultraviolet rays
 - d) γ -waves
16. A photon of energy 8 eV is incident on a metal surface of threshold frequency $1.6 \times 10^{15} \text{ Hz}$, then the maximum kinetic energy of photoelectrons emitted is ($h = 6.6 \times 10^{-34} \text{ Js}$)
 - a) 4.8 eV
 - b) 2.4 eV
 - c) 1.4 eV
 - d) 0.8 eV
17. The kinetic energy of an electron is 5 eV . Calculate the de-Broglie wavelength associated with it ($h = 6.6 \times 10^{-34} \text{ Js}$, $m_e = 9.1 \times 10^{-31} \text{ kg}$)
 - a) 5.47 \AA
 - b) 10.9 \AA
 - c) 2.7 \AA
 - d) None of these
18. Order of q/m ratio of proton, α -particle and electron is
 - a) $e > p > \alpha$
 - b) $p > \alpha > e$
 - c) $e > \alpha > p$
 - d) None of these
19. In the following diagrams if $V_2 > V_1$ then



- a) $\lambda_1 = \sqrt{\lambda_2}$
 - b) $\lambda_1 < \lambda_2$
 - c) $\lambda_1 = \lambda_2$
 - d) $\lambda_1 > \lambda_2$
20. Ultraviolet radiations of 6.2 eV falls on an aluminium surface. KE of fastest electron emitted is (work function = 4.2 eV)
 - a) $3.2 \times 10^{-21} \text{ J}$
 - b) $3.2 \times 10^{-19} \text{ J}$
 - c) $7 \times 10^{-25} \text{ J}$
 - d) $9 \times 10^{-32} \text{ J}$